

C8 Quick Revision Questions

H = Higher tier only

SS = Separate science only

Question 1

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State some processes of physical separation

Answer 1

.... of 50

- Filtering
- Crystallise
- Distillation
- Fractional distillation
- Chromatography

Question 2

.... of 50

Which has the higher boiling point, an impure or pure substance?

Answer 2

.... of 50

Impure substances have a higher boiling point

Question 3

.... of 50

Describe the melting point of an impure substance when compared to a pure substance

Answer 3

.... of 50

Impure substances have a lower melting point over a broader range of temperatures

Question 4

.... of 50

Define formulation

Answer 4

.... of 50

A mixture that has been designed as a useful product.

Question 5

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What is the equation to calculate the R_f values in chromatography?

Answer 5

.... of 50

$$R_f = \frac{\text{Distanced moved by the } \textit{substance}}{\text{Distanced moved by the } \textit{solvent}}$$

Question 6

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Describe the test for hydrogen gas

Answer 6

.... of 50

Holding a lit splint at the open end of the test tube. The hydrogen will burn and release a rapid 'pop' noise

Question 7

.... of 50

Describe the test for oxygen gas.

Answer 7

.... of 50

Place a glowing splint at the open end of the test tube. The splint should reignite.

Question 8

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Limewater is a common name for what substance?

Answer 8

.... of 50

Calcium hydroxide

Question 9

.... of 50

What happens to limewater if carbon dioxide is bubbled through it?

Answer 9

.... of 50

The limewater would become milky (cloudy)

Question 10

.... of 50

What happens to damp litmus paper when in contact with chlorine gas?

Answer 10

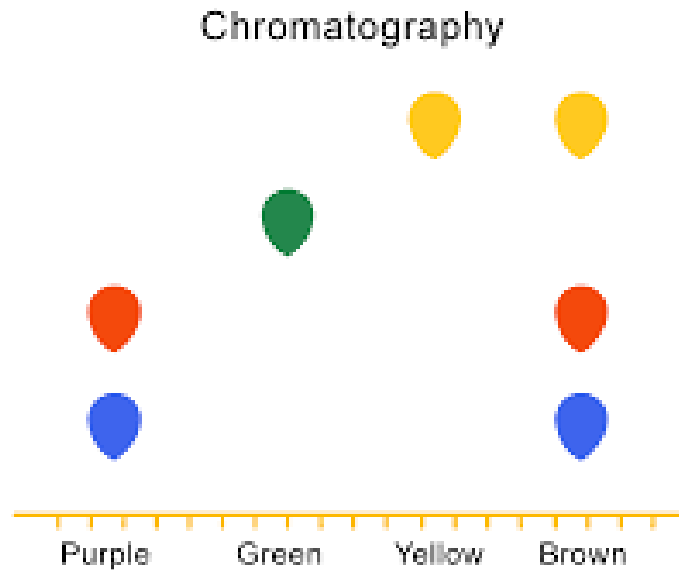
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The litmus paper is bleached and turns white

Question 11

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From this chromatogram, how many colours make up the purple dye?



Answer 11

.... of 50

two

SS

Question 12

.... of 50

What colour would the precipitate formed be, when sodium hydroxide is added to a solution containing iron(III) ions?

SS

Answer 12

.... of 50

Orange-brown

SS

Question 13

.... of 50

If sodium hydroxide is added to a solution containing iron (II) ions, what colour would the precipitate formed be?

SS

Answer 13

.... of 50

Green

SS

Question 14

.... of 50

What colour would the precipitate be when sodium hydroxide is added to solutions containing calcium, magnesium or aluminium ions?

SS

Answer 14

.... of 50

White

SS

Question 15

.... of 50

Silver nitrate is used to test for what ions?

SS

Answer 15

.... of 50

Halide ions (i.e. chloride, bromide, iodide)

SS

Question 16

.... of 50

Sulfate ions are tested for by using what substance?

SS

Answer 16

.... of 50

Barium chloride

SS

Question 17

.... of 50

Give three advantages of using instrumental techniques.

SS

Answer 17

.... of 50

More rapid

More accurate

More sensitive

SS

Question 18

.... of 50

What is flame emission spectroscopy used for?

SS

Answer 18

.... of 50

Analysing metal ions in solution

SS

Question 19

.... of 50

What colour would the flame be when potassium is burned?

SS

Answer 19

.... of 50

lilac

Which of the following ions would produce a yellow flame in a flame test?

- Lithium
- Sodium
- Potassium
- Calcium
- Copper

SS

Answer 20

.... of 50

Sodium

Which of the following ions would produce a green flame in a flame test?

- Lithium
- Sodium
- Potassium
- Calcium
- Copper

SS

Answer 21

.... of 50

Copper

